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Worksheet: Is Matter Around Us Pure ?

1. What is emulsion?
2. What is meant by a substance?
3. What is “solubility” of a solute?
4. To make a saturated solution, 36 g of sodium chloride is dissolved in 100 g of water at 293 K. Find its concentration at this temperature.
5. How would, you confirm that a colourless liquid given to you is pure water?
6. Write the steps you would use for making tea. Use the words, solution, solvent, solute, dissolve, soluble, insoluble, filtrate and residue.
7. Try segregating the things around you as pure substances or mixtures
8. Classify each of the following as a homogeneous or heterogeneous mixture: soda water, wood, air. soil, vinegar, filtered tea.
9. How will you separate a mixture containing kerosene and petrol (difference in their boiling points is more than 25°C), which are miscible with each other?
10. How are sol, solution and suspension different from each other?
11. What is Tyndall effect?
12. What is aerosol?
13. Why is crystallisation a better technique than the evaporation process?
14. What is crystallisation? Why is crystallisation used?
15. What type of mixtures are separated by the technique of crystallisation?
16. Differences Between Homogeneous and Heterogeneous Mixtures.
17. Why water is called universal solvent?
18. How can you convert saturated solution into unsaturated or vice-versa?
19. What is chromatography? What are its various applications and underline the basic principles involved?
20. State the properties of a suspension.